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Modern Chemistry Chapter 15  
Acid-Base Titration & pH Chapter

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15: Acid-Base Titration and pH Jeopardy Template. 1. When you conduct an acid-base titration, a. the pH of the solution must go up. b. the pH of the solution must go down. c. the pH of the solution must be 7.0 at the end point. d. the equivalence point must be

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Chapter 15 Acid Base Titration Ph  
Practice Test

Baylor, Scott / Chapter 15 Acids  
and Bases Chapter 15 Test Acids  
And Bases Answers when acids or

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Acids act as acids or bases, they form their conjugate and the strength is the inverse of their parent. and there are certain strong acids and strong bases. LiOH, NaOH, KOH - when act as bases (get rid of  $H^+$  ion) then left with Li Page 13/24

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Chapter 15 Test Acids Bases  
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ACIDS AND BASES 453 SECTION  
15-1 OBJECTIVES List five general  
properties of aqueous acids and  
bases. Name common binary

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Acids and oxyacids, given their chemical formulas. List five acids commonly used in industry and the laboratory, and give two properties of each. Define acid and base according to Arrhenius's theory of ionization. Explain the differences



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CHAPTER 15 Acids and Bases  
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3e (Tro). Chapter 15 Acids and  
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1) \_\_\_\_\_ is found in carbonated  
beverages due to the reaction of

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Carbon dioxide with water.

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the extra experience,  
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## Answers Answers

Chapter 15: Acids and Bases

Acids and Bases Arrhenius

Definitions: acids - compounds that produce an increase in  $[H^+]$  when dissolved in water bases - compounds that produce an increase in  $[OH^-]$  when dissolved

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in water Lewis Definitions: acids -  
electron pair acceptors bases -  
electron pair donors Brønsted-  
Lowry Definitions:

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Chapter 15 Acids Bases Review  
bases C. They can act as acids or  
bases when reacting with each  
other D. They can act as acids  
when reacting with HCl molecules  
14) 25.0 cm<sup>3</sup> of sulfuric acid  
solution reacts with 36.2 cm<sup>3</sup> of  
0.225 mol dm<sup>-3</sup> sodium

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hydroxide solution. The  
concentration of the Chapter 15  
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## Answers Answers

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Matching: 1. strong acids A. bitter  
2. pH B. ionize only slightly in  
solution 3. titration C. salt + H<sub>2</sub>  
(g) F 4. neutralization D. salt + ...

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and Bases. strong acid. weak  
acid. strong base. weak base. an  
acid that ionizes completely in a



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solvent. an acid that releases few hydrogen ions in aqueous solution. a base that ionizes completely in a solvent. a base that releases few hydroxide ions in aqueous solution.

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